Krypton

Krypton is obtained by fractional distillation of liquid air. Krypton is an inert gas and occurs in the air. It is used to fill vacuum tubes and light bulbs. Krypton is a chemical element and was discovered by the British chemist Sir William Ramsay in 1898. Ramsay discovered five of the noble gases. He extracted samples of most of the gases of air. He named them argon (from the Greek word for inactive), neon (new), xenon (stranger) and krypton (hidden). Helium (Sun) had been detected in the Sun's spectrum.

The noble gases are the elements that form group O of the periodic table. The six noble gases are so inert, or stable, that they rarely react with other materials. In chemistry, -noble- means unreactive. The noble gases have no smell, colour, taste and do not burn. The noble gases are, in order of lightest to heaviest: helium, neon, argon, krypton, xenon and radon. Radon is the heaviest of the noble gases, is radioactive and can cause lung cancer if inhaled in large enough quantities.

About 99% of our atmosphere is made up of nitrogen and oxygen. The gases in the remaining 1% are argon, carbon dioxide and tiny amounts of hydrogen, ozone, methane, carbon monoxide, helium, neon, krypton and xenon. The amount of water vapour depends on temperature and humidity.

Your task:

- Research the Scientific work Sir William Ramsay was involved in and write an essay. You may be able to locate some pictures on the Internet and print these. Present your research work on paper or to make it more challenging… design a web page. If you make a web page email the URL to http://teachingtreasures.com.au for link inclusion.

Atomic Number	2	10	18	36	54	86
Abbreviation	He	Ne	Ar	Kr	Xe	Rn
Atomic Weight	4.0	20.2	39.9	83.8	131.3	222
	helium	neon	argon	krypton	xenon	radon